**Percent Yield Lab**

**Problem**: Using a balanced equation, determine the limiting reactant and percent yield of the precipitate formed when 0.5 grams of lead (II) nitrate is mixed with 1.0 gram of potassium iodide.

**Pre-lab questions**:

1. What is a precipitate?
2. Write the reaction that occurs between your two reactants.
3. Determine which of the products is the precipitate.
4. What is a limiting reagent (or reactant)?
5. How do you determine the limiting reactant?
6. What is percent yield?
7. How is percent yield different from percent error?

**Data**:

* Record all information measured during the lab!

**Calculations**:

* Don’t forget to show all work!

**Analysis Questions**:

1. What is your limiting reagent?
2. How much lead (II) iodide should have been produced theoretically?
3. Calculate your percent yield.
4. List some sources of error that could have affected your results and explain HOW they would have affected the final answer.