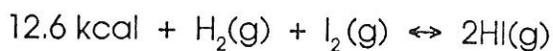


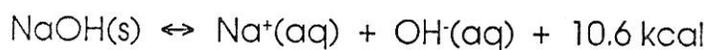
LE CHATELIER'S PRINCIPLE
CONTINUED

NAME _____

Key



Stress	Equilibrium Shift	[H ₂]	[I ₂]	[HI]	K
1. Add H ₂	right	—	decreases	increases	remains the same
2. Add I ₂	→	↓	—	↑	—
3. Add HI	←	↑	↑	—	—
4. Remove H ₂	←	—	↑	↓	—
5. Remove I ₂	←	↑	—	↓	—
6. Remove HI	→	↓	↓	—	—
7. Increase Temperature	→	↓	↓	↑	↑
8. Decrease Temperature	←	↑	↑	↓	↓
9. Increase Pressure	—	—	—	—	—
10. Decrease Pressure	—	—	—	—	—



(Remember that pure solids and liquids do not affect equilibrium values.)

Stress	Equilibrium Shift	Amount NaOH(s)	[Na ⁺]	[OH ⁻]	K
1. Add NaOH(s)	—	—	—	—	—
2. Add NaCl (Adds Na ⁺)	←	↑	—	↓	—
3. Add KOH (Adds OH ⁻)	←	↑	↓	—	—
4. Add H ⁺ (Removes OH ⁻)	→	↓	↑	—	—
5. Increase Temperature	←	↑	↓	↓	↓
6. Decrease Temperature	→	↓	↑	↑	↑
7. Increase Pressure	—	—	—	—	—
8. Decrease Pressure	—	—	—	—	—